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HOMEWORK 9-2 (pp. 280-282)

VOCABULARY ... Circle the letter of the best answer. 1. Given the balanced equation  $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$ , how many moles of  $\text{CO}_2$  are produced when 14.9 g of  $\text{O}_2$  ... Microsoft

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CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the  
answer on the line to the left. Show all

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your work in the space provided. 1. 88%

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $\text{N}_2$  are mixed with 12.0 mol of  $\text{H}_2$  according to the following equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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CHAPTER 6 REVIEW Chemical Bonding

SECTION 3 SHORT ANSWER Answer

the following questions in the space

provided. 1. a The notation for sodium chloride,  $\text{NaCl}$ , stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2.

d In a crystal of an ionic compound, each

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cation is surrounded by a number of (a)  
molecules. (c) dipoles. (b ...

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Stoichiometry. Sections 1-3 Introduction  
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Percent Yield. 2. Definitions. Composition  
stoichiometry deals with the mass.  
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