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required in terms of the relationship between temperature and pressure and the kinetic molecular theory.

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The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3

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motion of solids, liquids, and gases. 2.
KMT explains the differences between

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properties of solids, liquids, and gases
by examining how the particles are
moving

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Answer~~

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to predict the outcome of everyday situations. POGIL: Kinetic Molecular Theory The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the

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The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the

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following five postulates described here. (Note: The term “ molecule ” will be used to refer to the individual chemical species that compose the gas, although some ...

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The particles are assumed to not attract nor repel each other. The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3

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The Kinetic-Molecular Theory

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{Boyle 's Law zP 1/V zAs the V increases the molecular collisions with container walls decrease and the P decreases. {Dalton ' s Law zP total = P A + P B + P C + zBecause gases have few intermolecular attractions, their pressures are independent of other gases in the container.

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{Charles ' Law zV T

~~CHAPTER 12 GASES AND KINETIC- MOLECULAR THEORY~~

The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3

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Linda Padwa and David Hanson, Stony Brook University Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? walls 2.

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The observed behavior of gases, as expressed by the empirical gas laws, can be understood on the basis of the kinetic molecular theory, which developed over many years up to 1857, when Rudolf Clausius (1822-1888) published it in its most complete and elegant form.

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Studyres The Kinetic-Molecular
Theory Explains the Behavior of
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law, the molecules of a gas are in

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rapid motion and the molecules themselves are small. The average distance between the molecules of a gas is large compared to the size of the molecules. Page 1/5

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Answer - Bespokify~~

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Kinetic Molecular Theory states that gas particles are in constant motion and exhibit perfectly elastic collisions. Kinetic Molecular Theory can be used to explain both Charles ' and Boyle ' s Laws. The average kinetic energy of a collection of gas particles is directly proportional to absolute

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temperature only.

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theory is the theory that explains the
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Kinetic Molecular Theory is guided by

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the following assumptions: The molecules of an ideal gas (an ideal gas is a gas that follows the assumptions of kinetic molecular theory) are in constant, random, straight-line motion. Gas molecules are constantly moving very, very fast and their direction is completely random.

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kinetic molecular theory answer key,
Feb 19, 2003 · 2. State the six
postulates of the Kinetic-molecular
theory and explain how they account
for the physical properties of gases.

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Key Terms: diffusion elastic Kinetic-
molecular theory. Notes: (13-1)
Nature of Gases: We are constantly
surrounded and buried in gas
molecules.

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