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quantity to decay to one-half of its original amount. Since $\lambda = 1/\tau$, we get $\frac{1}{2} r_0 = r_0 e^{-\lambda h}$ $\frac{1}{2} r_0 = r_0 e^{-h/\tau}$ $\frac{1}{2} = e^{-h/\tau}$ $-\ln 2 = -h/\tau$. Thus, $\tau = h \ln 2$ (c) In Exercise 7, we stated that the half-life of Carbon 14 is 5230 ...

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