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Calculating pH from

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~~pH and pOH |~~

~~Chapter: Acids, Bases
and Salts | Class 10th~~

~~Chemistry 10th Class~~

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~~\u0026 POH - Matric~~

~~Part 2 Chemistry pH~~

~~scale/ relationship~~

~~between pH and pOH/~~

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~~Calculating pH, pOH,~~

~~[H⁺], [H₃O⁺], [OH⁻]~~

~~of Acids and Bases -~~

~~Practice pH of Weak~~

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Acids and Bases, Salt
Solutions, K_a , K_b , pOH
Calculations

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~~Concept of pH and
 pOH Chemistry Ph And
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~~Gridlocks - Can you
unlock the grid? | pH
and pOH~~

$$pH = -\log[H^+] = -\log(1.0 \times 10^{-7}) = 7.0$$

$$pOH = -\log[OH^-] = -\log(1.0 \times 10^{-7}) = 7.0$$

Here we have the reason that neutral water has a pH of 7.0 -; this is the pH at which the concentrations of H

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exactly equal. Lastly, we should take note of the following relationship: $pH + pOH = 14$

~~pOH and Other p
Scales | Introduction to
Chemistry~~

Chemistry Ph And Poh
Grid Answers $pH =$
 $14.000 - pOH =$
 $14.000 - 2.187 =$

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11.813 4) -A solution is created by measuring 3.60×10^3 moles of NaOH and 5.95×10^{-4} moles of HCl into a container and then water is

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Grid Answers pH =

$$14.000 - \text{pOH} =$$

$$14.000 - 2.187 =$$

11.813 4) -A solution is

created by measuring

3.60×10^{-3} moles of

NaOH and 5.95×10^{-4}

moles of HCl into a

container and then

water is added until the

final volume is 1.00 L.

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Chemistry Ph And Poh
Grid Answers pH =

$$14.000 - \text{pOH} =$$

$$14.000 - 2.187 =$$

11.813 4) -A solution is

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NaOH and 5.95×10^{-4}

moles of HCl into a

container and then

water is

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pH 1; pH and pOH.

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different indicators react
with acids and ...

~~Gridlocks Can you
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This lesson introduces
the pH scale and
discusses the
relationship between
pH, $[H^+]$, $[OH^-]$ and

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~~Chemistry 12.3a pH
and pOH - YouTube~~

There are a few different formulas you can use to calculate pOH, the hydroxide ion concentration, or the pH (if you know pOH):
$$\text{pOH} = -\log_{10} [\text{OH}^-]$$
$$[\text{OH}^-] = 10^{-\text{pOH}}$$
$$\text{pOH} + \text{pH} = 14 \text{ for any aqueous solution.}$$

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$$\text{pH} = -\log [\text{H}^+]$$
$$= -\log (4.9 \times 10^{-7})$$

$$= 6.31. \text{pOH} = -\log$$
$$[\text{OH}^-] = -\log (4.9$$

$\times 10^{-7}) = 6.31.$ At
this temperature, then,
neutral solutions exhibit
 $\text{pH} = \text{pOH} = 6.31,$
acidic solutions exhibit

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pH less than 6.31 and
pOH greater than 6.31,
whereas basic solutions
exhibit pH greater than
6.31 and pOH less than
6.31.

~~14.2 pH and pOH |
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The pH and pOH of a solution are related such that: $\text{pH} + \text{pOH} = 14$. For example, if the pH of a solution is 5, the pOH of the same solution will be $14 - 5 = 9$. The pH of distilled water is 7, this is neutral. Any solution with a pH below 7 (i.e. pH 0 to pH 6.9) is acidic and any solution with a

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pH above 7 (i.e. pH 7.1 to pH 14) is basic.

Structural

~~Biochemistry/pH~~

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for an ...~~

pH is a figure expressing the acidity or alkalinity of a solution on a logarithmic scale on which 7 is neutral. The values lower than 7 are more acidic while higher

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values more alkaline.

The pH is equal to

$-\log_{10} c$, where c is the hydrogen ion concentration in moles per litre. The pH scale runs from 1 to 14.

~~Difference Between pH and pOH | Compare the Difference ...~~

9.4 pH and pOH Last updated; Save as PDF
Page ID 218433; Strong

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Acids; Strong Bases;
Weak Acids and Weak
Bases; If we wish to find
the hydronium ion
concentration ($[H_3O^+]$) and the pH of a
solution, we need to
know both the strength
of the acid (or base) and
the concentration of the
acid (or base). We will
find that we need to
treat strong acids (and
bases) differently than

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Answers

~~9.4 pH and pOH~~

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~~Indicators~~

pOH = 14 - pH While
the approximation
works well in many

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settings, there are exceptions for which the pK w value should be used instead. Here's
How to Calculate pH
Values

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