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unit describing CONCENTRATION of  
a solution. solutions. homogeneous  
mixture of 2 or more pure substances.  
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solvent. in greater quantity than  
solute, typically a liquid but not always.  
solute.

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homogeneous mixture. The first step

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of solvation is to \_\_\_\_\_ thermal energy, whereas the mixing process \_\_\_\_\_ it. to absorb; release. A student desires to remove the color of naphthalene (moth balls), a nonpolar aromatic hydrocarbon, from his laboratory coat. What solvent should he ...

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Classify the following bonds as  
nonpolar covalent, polar covalent, or

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ionic. If a bond is polar covalent, identify which atom has the partial negative charge and which has the partial positive charge. If a bond is ionic, identify which atom

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capacity of 1 gram of a substance. the  
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small b. forces c. random d. elastic;  
kinetic 3.  $KE = \frac{1}{2}mv^2$  4. Temperature  
5. true 6. true 7. false 8. true 9. true  
10. false 11. true 12. false 13. a 14. a  
15. d 16. d 17. b 18. b ...

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30. Compare and contrast sublimation and evaporation. In both processes, the substances become a vapor.

During sublimation, the substance



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goes from the solid phase directly to the vapor phase. During evaporation, particles in a liquid gain enough

*States of Matter*

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textbook, read about mole-to-mole conversion. Read the following passage and then solve the problems. In the equation that follows each problem, write in the space provided the mole ratio that can be used to solve the problem.

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Aldehydes, Ketones and Carboxylic  
Acids. The next ones in line are the  
aldehydes, ketones and carboxylic  
acids. The IUPAC names, to start with,  
the chapter explores their preparation

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methods, the key reactions,  
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categories of science. the science of  
different kinds of matter and how that  
matter can change. Define Chemistry.  
organic. the study of carbon based

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