

## Chemistry 2nd Semester Study Guide Part 1

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1) calculate moles of each element (if given mass) 2) divide each mole quantity by the smallest mole value from step 1. 3) use whole number values from step 2 as the subscripts to the original equation to write the empirical formula. Percent Composition. -the relative amounts of each element in a compound.

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Chemistry - 2nd Semester Practice Test Thermochemistry (Ch.11): Multiple Choice 24) What is the amount of heat required to raise the temperature of 200.0 g of aluminum by 10 °C? cal (specific heat of aluminum = 0.21 J/g °C) A) 4200 cal B) 420 cal C) 42 000 cal D) 420 000 cal E) none of the above

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2nd Semester Chemistry-575 Final Exam Review Multiple Choice Identify the choice that best completes the statement or answers the question. \_\_\_\_ 1. How many moles of tungsten atoms are in  $4.8 \times 10^{25}$  atoms of tungsten? a.  $8.0 \times 10^2$  moles c.  $1.3 \times 10^{-1}$  moles b.  $8.0 \times 10^1$  moles d.  $1.3 \times 10^{-2}$  moles \_\_\_\_ 2.

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