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You are already familiar with the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you. For example, you can easily recognize the shape of your desk; you know that your backpack cannot hold seven textbooks. You encounter

liquids throughout the day as yo u

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applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385 – 386) 1.

Name Date Class STATES OF MATTER 13

There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter. I can refer you to somebody.) Those three states are solid, liquid, and Page 16/36

gas. These three states are quite different. The main difference is in their particles.

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kinetic energy of the particles of a substance.

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gases in terms of the energy, size, and motion of their particles. Elastic Collision, Describes a collision in which kinetic energy may be transferred between colliding particles but the total kinetic energy of the two particles remains the same.

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all matter consists of tiny particles that are constantly in motion What are the three assumptions of the kinetic theory as it applies to gases? - The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the Page 26/36

particles in a gas are rapid, constant, and random.

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The Nature of Solids F13.3 Lesson Check

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Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

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