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Chapter 11

Chapter 11 S

toichiometry

Study Guide

Answer Key

Study Guide for  
Chemical Principles  
[by] Steven S.  
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Chemistry Study  
Guide & Solutions

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Stoichiometry  
Practice Problems |  
How to Pass  
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Chemistry: Chapter

11 \u0026 12 Unit

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General Chemistry

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Guide - IB, AP,

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Chem Final Exam

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CONCEPT for JEE?

(by a JEE topper)

Know This For Your  
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Chemistry:  
chemical reactions  
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Basic Concepts of  
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11 | in Hindi Chapter

11 Stoichiometry

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Chapter 11 -

Stoichiometry

Section 11.1 What

is stoichiometry? 1.

true 2. true 3. false

4. true 5. true 6. 2,

2, 64.10 7. 3, 3,

96.00 8. 2, 2, 88.02



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9. 4, 4, 72.08 10.

methanol and  
oxygen gas 11.

carbon dioxide and  
water 12. 160.10 g

13. 160.10 g 14.

They are equal. 15.

A mole ratio is a  
ratio between the  
numbers of moles

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Technology

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368 Chapter 11 □

Stoichiometry

Section 11.11.1

Objectives

Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary  
reactant: the

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Starting substance  
in a chemical  
reaction New  
Vocabulary

stoichiometry mole  
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Concepts Using

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□ Number of Atoms

□ Number of

Molecules □

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Number of Moles □

Mass o Law of  
Conservation of  
Mass □ Volume

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Stoichiometry ...

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Stoichiometry

Section 11.1 What is stoichiometry? In your textbook, read about

stoichiometry and the balanced equation. For each statement below, write true or false.

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### 2.3.4.5. The

study of the  
quantitative  
relationships

between the  
amounts of  
reactants used and  
the amounts of  
products formed by  
a chemical reaction

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□ Number of Atoms

□ Number of

Molecules □

Number of Moles

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### Stoichiometry

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of

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the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a

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reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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