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the virtual lab activity. Record your answers in the boxes.

Reference: Isotope Half-Life Chart Isotope Product Half Life

Carbon-14 Nitrogen-14 5730 years Potassium - 40 Argon - 40

1,280 million years Rubidium - 87 Strontium -87 48,800

million years Thorium - 232 Lead - 208 14,010 million years

Uranium - 235 Lead - 297 704 million years Uranium - 238

Lead - 206 4,470 million ...

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Calculate the age of the calibration standards using the following information. Fraction of sample remaining = remaining ppm of sample/initial ppm of sample Age of sample = half-life value of isotope X number of half-lives elapsed

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Calibration Standard Initial ppm Remaining ppm Age of  
Standard Low Carbon-14 12000 ppm 5998 ppm 5730 years  
High Uranium- 235 600000 pp, 151000 ppm 1408 million  
years

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Uranium - 235 Lead - 297 704 million years Uranium - 238

Lead - 206 4,470 million years Activity 1 – Calibration Place

your data from Activity 1 in the appropriate boxes below.

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Place your data from Activity 1 in the appropriate boxes below. Calculate the age of the calibration standards using the following information. \* Fraction of sample remaining = remaining ppm of sample/initial ppm of sample \* Age of

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sample = half-life value of isotope X number of half-lives elapsed.

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Isotope Product Half Life  
Carbon-14 Nitrogen-14 5730 years  
Potassium - 40 Argon - 40 1,280 million years  
Rubidium - 87

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This is the fundamental age equation, where  $P$  is the number of atoms of a given radioactive element at time  $t$ , and at time  $t+0$ , the original number of atoms is  $P_0$ . Derivation of the Half Life Equation. To calculate the half life we set  $P$  now equal to half of the original sample, which yields:  $0.5 P_0 = P_0 e^{-\lambda t}$ .

## Some common terms - Columbia University

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the following questions about the results of this activity.

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remaining ppm of sample/initial ppm of sample. Age of sample = half-life value of isotope X number of half-lives elapsed. Calibration Standard.

## Lab Report - Wiley

Carbon-14 ( $^{14}\text{C}$ ), or radiocarbon, is a radioactive isotope of carbon with an atomic nucleus containing 6 protons and 8 neutrons. Its presence in organic materials is the basis of the radiocarbon dating method pioneered by Willard Libby and colleagues (1949) to date archaeological, geological and hydrogeological samples. Carbon-14 was discovered on February 27, 1940, by Martin Kamen and Sam ...

## Carbon-14 - Wikipedia

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Principles of isotopic dating All absolute isotopic ages are based on radioactive decay, a process whereby a specific atom or isotope is converted into another specific atom or isotope at a constant and known rate.

Dating - Principles of isotopic dating | Britannica

post lab bean bag isotopes simple, it was revolutionary. Scientists became able to study the actual structure and mass of atoms after the discovery of radioactivity. Soon, isotopes were discovered, as atoms of the same element which have been built up to have different masses.

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