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#27 Acids, Bases

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Acids And Bases

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Answer Key
Key is an
example of a

strong acid.

Sodium hydroxide

(NaOH) is an

example of a

strong base. 8.

According to the

definitions of

acids and bases

devised by

Lewis: acids are

electron pair

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acceptors and
bases are
electron pair
donors. acids
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pair donors and

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Acids and bases

interact with

each other in

what is called a

neutralization

reaction. The

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Products of the reaction are a salt and water.

The products of the reaction are a salt and water. The pH is "neutralized" to 7 if both the acid and base fully react.

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Acids And Bases

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Bases And Salts

Guided A

substance is

called base if

its aqueous

solution tastes

bitter, turns

red litmus blue

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or neutralizes acids. Salt is a neutral

substance whose aqueous solution does not affect litmus.

According to Faraday: acids, bases, and salts are termed as electrolytes.

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Brønsted-Lowry

Acids & Bases

Identify each

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Acids And

Species in the

following

equation as

either the

Brønsted-Lowry

acid, the

Brønsted-Lowry

base, the

conjugate acid,

or the conjugate

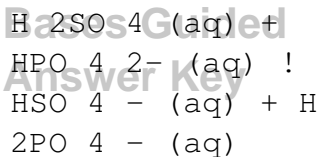
base. Identify

the conjugate

acid-base pairs

in the reaction.

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Strong vs. Weak
Acids and Bases
Strong Acids:

Chapter 15:
Acids and Bases
Acids and Bases
Acids and bases
can be defined
via three

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different theories. The

Arrhenius theory

of acids and

bases states

that “an acid

generates H^+

ions in a

solution whereas

a base produces

an OH^- ion in

its solution”.

The Bronsted-

Lowry theory

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defines “an acid as a proton donor and a base as a proton acceptor” .

*Acids and Bases
- Definition,
Examples,
Properties, Uses
...*

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Answer Key

Acids and Bases

An acid is any substance whose aqueous solution is characterized

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by a sour taste,
the ability to
turn blue litmus
red, and the
ability to react
with bases and
certain metals
to form...

*Answers about
Acids and Bases*

In water or a
water solution,
the solution is

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acidic if the hydrogen ion concentration is greater than the hydroxide (OH^-) ion

concentration, the solution is basic if the hydroxide ion concentration is greater than the hydrogen (H^+) ion

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Bases Concentration,
and the solution
is neutral when
the
concentrations
are equal.

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bases and salts
guided answer
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that this is the
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you are in the
bus, office,
home, and
supplementary
places. But, you

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may not
dependence to
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involve or bring
the cassette
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Answer Key

This becomes
important in the
Brønsted-Lowry
definition of
acids and bases

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Bases Guided

because acids
are chemicals
that donate

protons and

bases are ones

that accept

protons. I then

pass out the

Acids and Bases

Reading

Questions which

guide students

in finding the

background

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knowledge I want
students to
learn for this
unit from their
text. I note
that they should
work on
questions 1-19
and also work
with a partner
to study their
acid and base
names.

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Lesson

Answer Key

Introduction to

Acid Base Theory

Acids and Bases

Acid and Base

Strength In any

acid-base

reaction, the

equilibrium will

favor the

reaction that

moves the proton

to the stronger

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Acids And

base. $\text{HCl} (\text{aq})$

$+ \text{H}_2\text{O} (\text{l}) \rightarrow$

$\text{H}_3\text{O}^+ (\text{aq}) + \text{Cl}^-$

(aq)

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A substance is

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called base if
its aqueous
solution tastes

bitter, turns

red litmus blue

or neutralizes

acids. Acids

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base strength is

based on the

extent of

ionization that

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Acids And

Describing Acids

and Bases Review

and Reinforce 1.

Sour 2. Bitter

3. Corrosive to
magnesium, zinc,

and iron; eats

them away and

produces bubbles

of hydrogen gas

4. Doesn't react

with metals 5.

Produces carbon

dioxide 6.

