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answer on the line to the left. Show all  
your work in the space provided. 1.

88% The actual yield of a reaction is  
22 g and the theoretical yield is 25 g.

Calculate the percentage yield. 2. 6.0

mol of  $N_2$  are mixed with 12.0 mol of  
 $H_2$  according to the following

equation:  $N_2(g) + 3H_2(g) \rightarrow$  ...

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chapter 9 - stoichiometry - yazvac labs

in chapter 9 include a study of

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