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## 22 1 Nuclear Chemistry Answer Key

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~~Particles, Beta Particles, Gamma Rays, Positrons, Electrons,~~  
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represented Define and relate the terms mass defect and nuclear binding energy 22 1 Nuclear

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Table 22.1 and example 22.1, problems 22.1-2 Beta emission (occurs when an atom has too many neutrons, or not enough protons). In the nucleus a neutron splits into an electron and a proton ( $1\ 0\ n \rightarrow 1\ 1\ p + 0\ -1\ e$ ) Since this electron originated from the nucleus, it is a beta particle.

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natural, humans can't control it, stop it, or slow it down.

Transmutation: when atoms are hit or bombarded by particles a reaction may occur. Chapter 22 Review Nuclear Chemistry Chapter 22 nuclear chemistry. Mass defect.

Nucleons. Nuclide. Radioactive decay. The difference in mass of an atom and Page 11/27

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Chapter 21 - Nuclear Chemistry - Questions and Exercises ... Example of nuclear fusion reaction is  $1\text{ }^2_1\text{H} + 1\text{ }^2_1\text{H} \rightarrow 2\text{ }^4_2\text{He} + 23.84\text{MeV}$ . Nuclear fusion reaction is used in making hydrogen bombs. 5) Write nuclear reactions for the following transformation: a) Emission of  $\alpha$ -particle from  $^{238}_{92}\text{U}$ . Ans-  $^{238}_{92}\text{U} \rightarrow \alpha + ^{234}_{90}\text{Th}$ . b) Emission of positron  $\beta^+$  from  $^{27}_{14}\text{Si}$ .

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