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techniques (ie, detect odor, light

burner, read volume, find mass,

etc.) phases/states -

characteristics. density formula - be able to calculate.

percentage/relative error formula

- be able to calculate. evidence of

reactions. ways to start reactions.

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1. form solids in room temp.

(crystalline units) 2. brittle 3. high melting points (1000°C or higher)

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4. conduct electricity in molten state (electrolytes- charged particles that conduct electricity)
5. dissolves in polar solvents (dissociation- water)
6. best way to test for ionic compound - test w/ electrical conductivity in a solution

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Semester 1: Chemistry Final Study Guide Chemistry semester 1 exam study guide. 4. Complete the Table: Periodic Trend
Direction Left to Right
Electronegativity Increase or Decrease Increase The Reason If proton number increases, then positive charge of nucleus increases. If there is more

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positive charge, then the nucleus will attract. . .

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1. Name the compound with the formula $\text{Ni}_2(\text{CO}_3)_3$. 2. Write the formula for aluminum sulfate. 3. Name the compound with the formula N_2O_5 . 4. Write the formula for phosphorus trichloride. 5. Name the compound with the formula H_3PO_4 . 6. Write the formula for hydrosulfuric acid. Chapter 8 Molecular Shape. Summary. 1. What is the VSEPR theory? 2.

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