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(ie, detect odor, light burner, read volume, find mass, etc.)

phases/states - characteristics. density formula - be able to calculate.

percentage/relative error formula - be able to calculate. evidence of

reactions. ways to start reactions.

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1. form solids in room temp. (crystalline units) 2. brittle 3. high melting

points (1000 ° C or higher) 4. conduct electricity in molten state

(electrolytes- charged particles that conduct electricity) 5. dissolves in

polar solvents (dissociation- water) 6. best way to test for ionic

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compound - test w/ electrical conductivity in a solution

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1. Name the compound with the formula  $\text{Ni}_2(\text{CO}_3)_3$ . 2. Write the formula for aluminum sulfate. 3. Name the compound with the formula  $\text{N}_2\text{O}_5$ . 4. Write the formula for phosphorus trichloride. 5. Name the compound with the formula  $\text{H}_3\text{PO}_4$ . 6. Write the formula for hydrosulfuric acid. Chapter 8 Molecular Shape. Summary. 1. What is the VSEPR theory? 2.

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